



UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS
International General Certificate of Secondary Education

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CO-ORDINATED SCIENCES (US)

0442/13

Paper 1 Multiple Choice

October/November 2013

45 minutes

Additional Materials: Multiple Choice Answer Sheet
Soft clean eraser
Soft pencil (type B or HB is recommended)

* 9 8 0 8 2 1 8 2 5 2 *

READ THESE INSTRUCTIONS FIRST

Write in soft pencil.

Do not use staples, paper clips, highlighters, glue or correction fluid.

Write your name, Center number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

DO NOT WRITE IN ANY BARCODES.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A, B, C** and **D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

Any rough working should be done in this booklet.

A copy of the Periodic Table is printed on page 20.

Electronic calculators may be used.

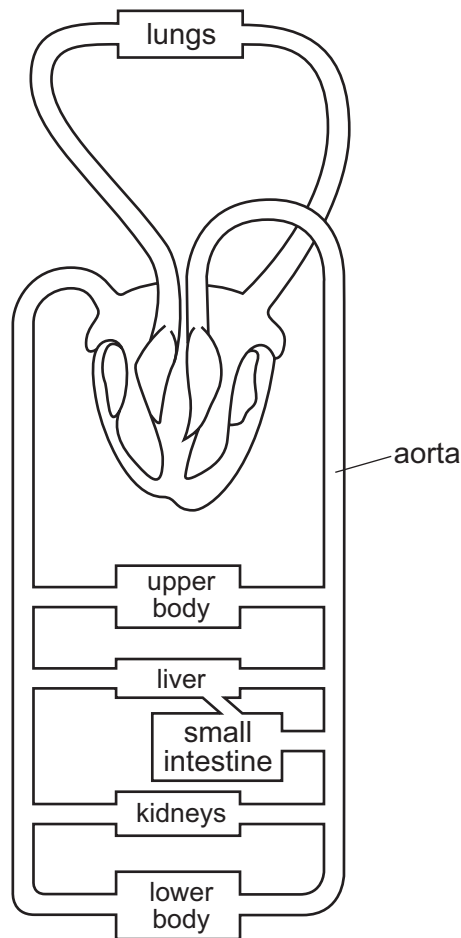
This document consists of **19** printed pages and **1** blank page.



- 1 A student was walking through some grass when he saw an object with a hole in its top surface and a small cap covering. When touched, a jet of fluid came out of the hole and the object moved away.

Which characteristics of living organisms has the object shown?

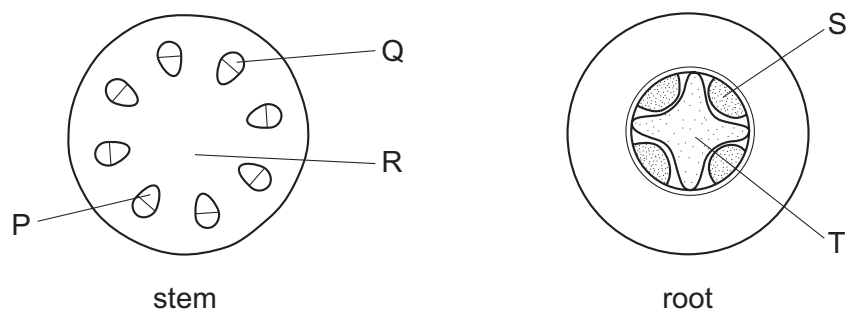
- A excretion, nutrition
 - B movement, respiration
 - C movement, sensitivity
 - D nutrition, sensitivity
- 2 The diagram shows the blood circulatory system of a human.



How many times must a blood cell pass through the heart on its way from the kidneys to the aorta?

- A once only
- B twice only
- C four times
- D more than four times

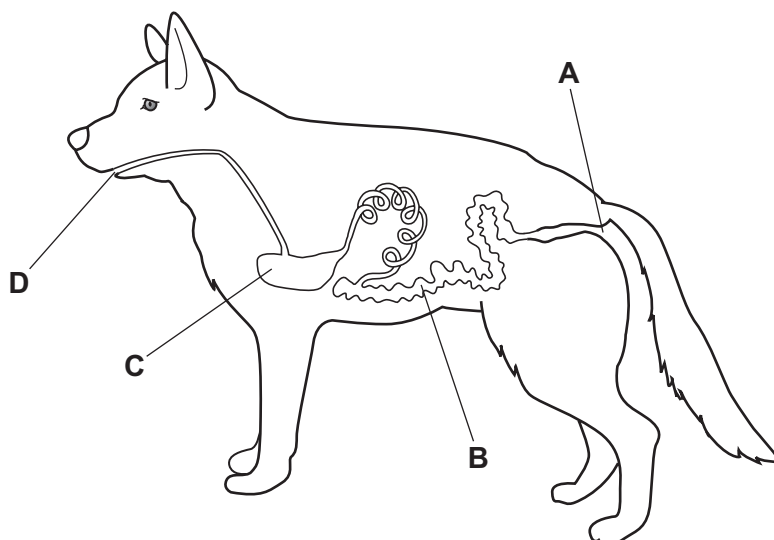
- 3 The diagrams show sections through a stem and a root.



Which indicate the positions of the phloem?

- A** P and S **B** P and T **C** Q and S **D** R and T
- 4 The diagram shows the alimentary canal of a dog.

Where does egestion occur?



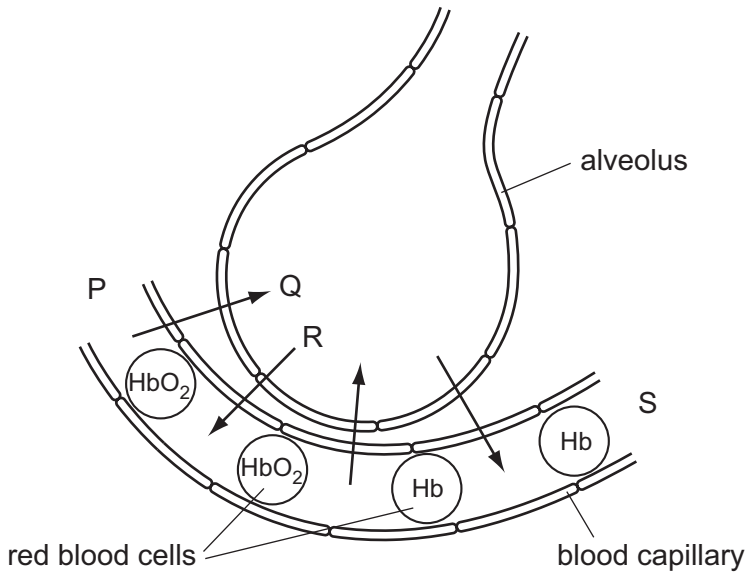
- 5 Which statement about asexual reproduction is correct?

- A** It involves the formation of a haploid zygote.
B It involves the fusion of haploid nuclei.
C It produces offspring that are genetically dissimilar to their parents.
D It produces offspring that are genetically identical to one another.

6 Which structures make up the nervous system?

- A brain, nerves, spinal cord
- B effectors, impulses, spinal cord
- C impulses, muscles, nerves
- D effectors, receptors, stimuli

7 The diagram shows an alveolus, a blood capillary and some red blood cells.



key

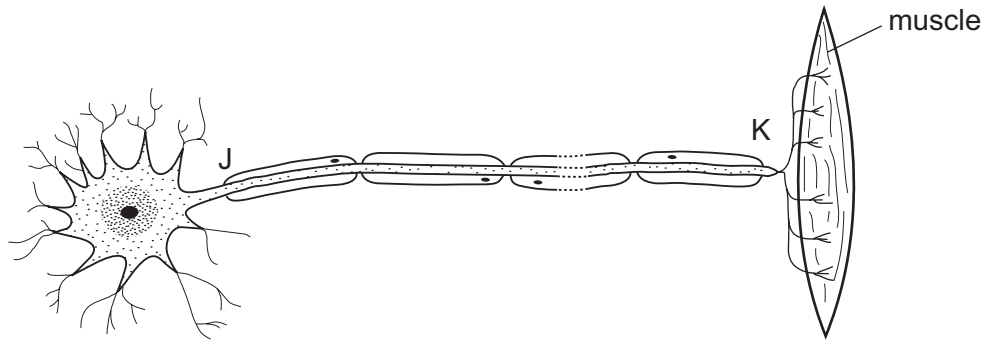
Hb = hemoglobin

HbO₂ = hemoglobin combined with oxygen

What is the direction of blood flow in the capillary and the direction of diffusion of oxygen?

	blood flow	oxygen diffusion
A	P to S	Q
B	P to S	R
C	S to P	Q
D	S to P	R

- 8 The diagram shows a nerve cell and associated structures.

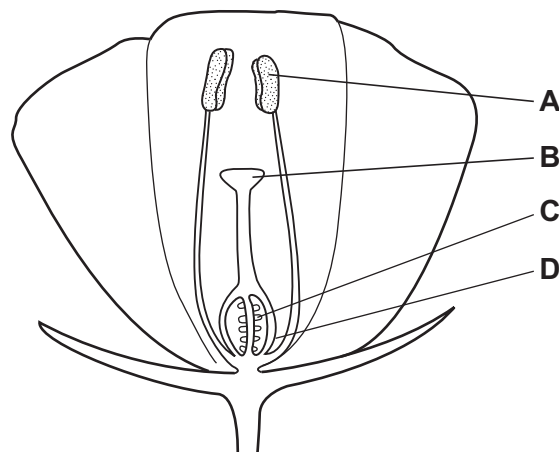


What type of nerve cell is it and in which direction do impulses travel?

	type of nerve cell	direction of impulse
A	motor	J to K
B	motor	K to J
C	sensory	J to K
D	sensory	K to J

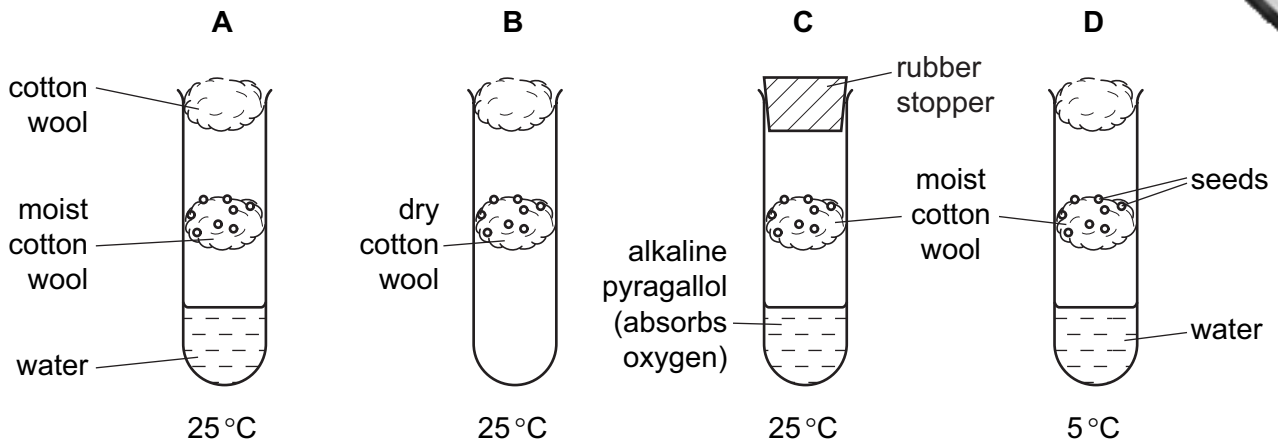
- 9 The diagram shows a section through an insect-pollinated flower.

When pollination occurs where must the pollen grains reach?



- 10 Seeds were placed on cotton wool in each of the tubes shown in the diagrams.

In which tube would germination start first?



- 11 The alleles for a particular character are H and h.

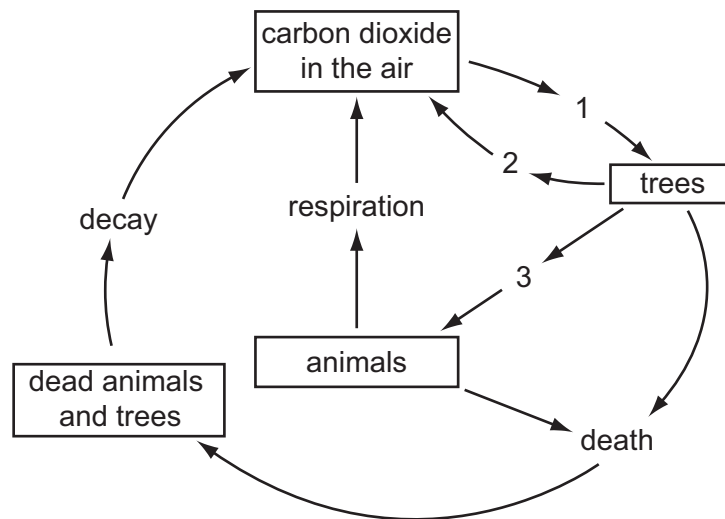
Which term describes an organism whose genotype is HH?

- A heterozygous
 - B homozygous
 - C phenotype
 - D recessive
- 12 Dung beetles lay their eggs in the faeces of plant-eating mammals like buffalo. Both the adult beetles and their young stages eat the **undigested** food in the faeces.

Which shows this food relationship?

- A buffalo → dung beetles
buffalo → grass
- B dung beetles → grass → buffalo
- C grass → dung beetles → buffalo
- D grass → buffalo
grass → dung beetles

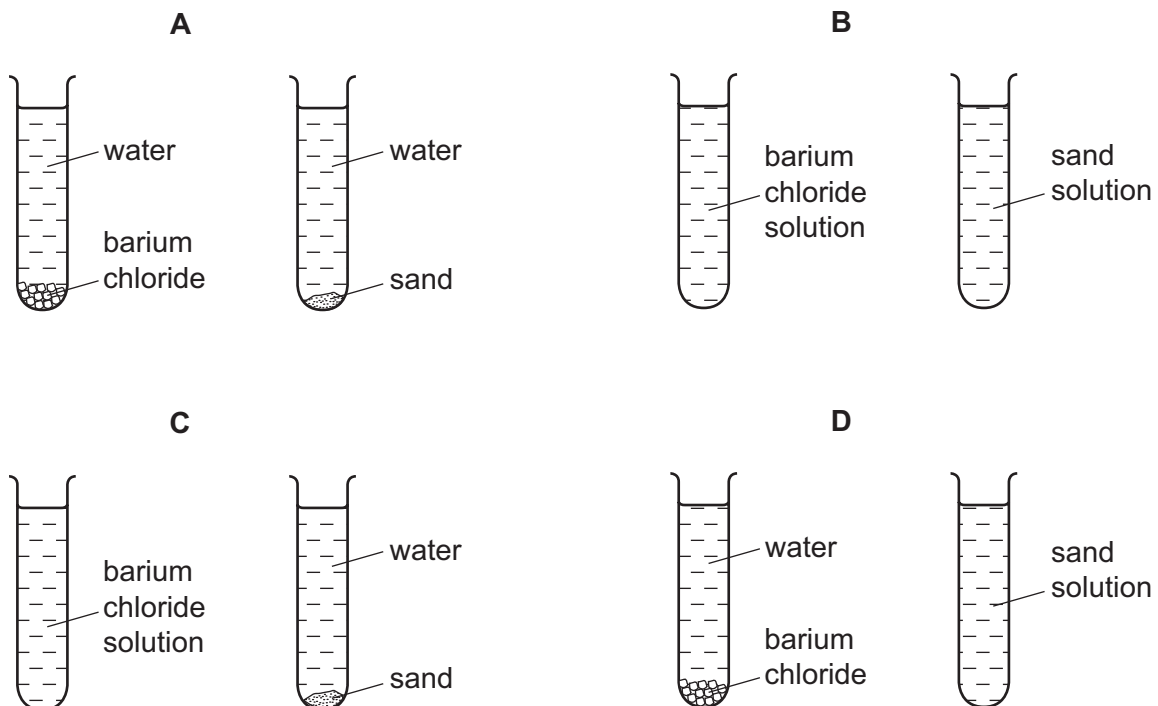
- 13 The diagram shows part of the carbon cycle in a forest. The numbers represent processes.



Which of these processes is reduced as a result of deforestation?

- A 1 only
 B 1 and 2 only
 C 2 and 3 only
 D 1, 2 and 3
- 14 Small amounts of barium chloride and sand are shaken with separate samples of water in two test-tubes. The test-tubes are left to stand for 24 hours.

Which diagram shows how the test-tubes appear at the end?



15 Which of the substances can conduct electricity?

	solid copper	molten copper	solid sodium chloride	molten sodium chloride
A	✓	✓	✓	✓
B	✓	✓	✗	✓
C	✗	✓	✓	✓
D	✗	✓	✗	✓

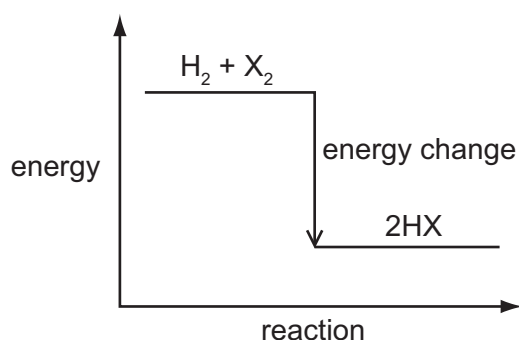
key

✓ = conduct

✗ = does not conduct

16 The diagram shows the energy change for the reactions between hydrogen and the halogens.

The size of the energy change is different for each halogen.



The reaction is $\text{H}_2 + \text{X}_2 \rightarrow 2\text{HX}$

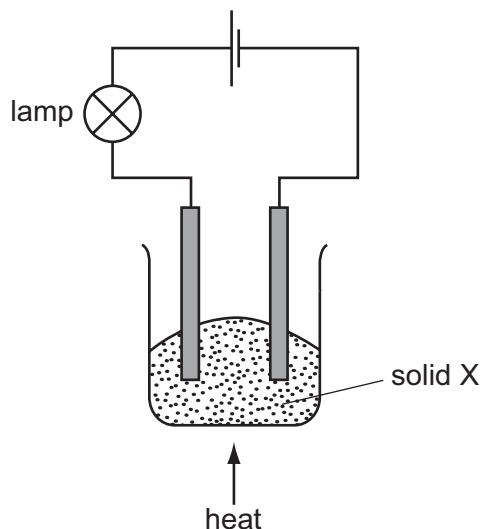
The diagram shows that the reactions are1..... .

The most reactive halogen is2..... and therefore the energy change for this element is3..... .

Which words complete gaps 1, 2 and 3?

	1	2	3
A	endothermic	fluorine	least
B	endothermic	iodine	least
C	exothermic	fluorine	greatest
D	exothermic	iodine	greatest

- 17 The experiment shown is used to investigate the properties of solid X.

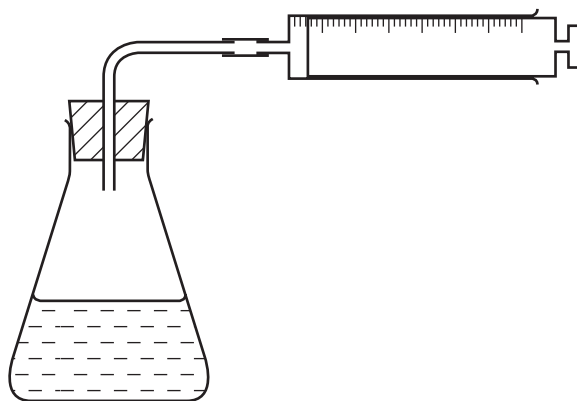


At first, the lamp does not light.

On heating, solid X melts and the lamp lights.

What type of substance is X?

- A a compound of a metal and a nonmetal
 - B a compound of two nonmetals
 - C a metallic element
 - D a nonmetallic element
- 18 The diagram shows apparatus used to investigate the speed of a reaction.



Which other item is essential for this investigation?

- A a Bunsen burner
- B a graduated cylinder
- C a stopclock
- D a thermometer

- 19 Brine is a mixture of salt (sodium chloride) and water.

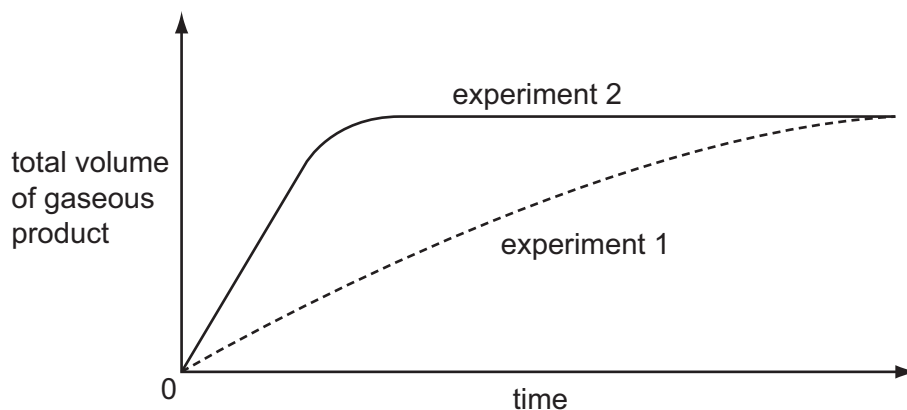
Which row describes these substances?

	solute	solvent	solution
A	brine	salt	water
B	brine	water	salt
C	salt	brine	water
D	salt	water	brine

- 20 Substance X does not react with dilute acid. Substance Y reacts with dilute acid, forming a gas.

The graph shows the results of two experiments.

experiment 1 Y + dilute acid
experiment 2 X + Y + dilute acid



What do these results show?

	X is a catalyst	X is quickly used up
A	✓	✓
B	✓	✗
C	✗	✓
D	✗	✗

key

✓ = true

✗ = false

- 21 The box shows four substances.

Br ₂	CO	Cu	Na
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Which substance is an element that forms a basic oxide and colored compounds?

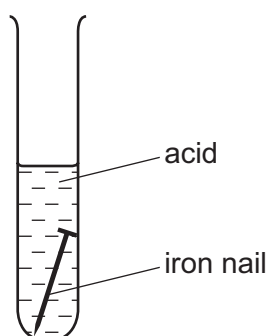
- A** Br₂ **B** CO **C** Cu **D** Na

- 22 A cup is made of copper.

Why is the cup **not** used for hot drinks?

- A Copper is a good conductor of heat.
- B Copper is a good electrical conductor.
- C Copper is brightly colored.
- D Copper reacts with saliva.

- 23 An iron nail dissolves in an acid to form a salt solution.



The salt solution forms a green precipitate with sodium hydroxide solution.

The salt solution also forms a white precipitate with barium chloride solution.

What is the salt solution?

- A iron(II) chloride
 - B iron(III) chloride
 - C iron(II) sulfate
 - D iron(III) sulfate
- 24 Which type of reaction and which temperature change take place when an acid reacts with an alkali?

	type of reaction	temperature change
A	endothermic	decrease
B	endothermic	increase
C	exothermic	decrease
D	exothermic	increase

25 The elements in a Group of the Periodic Table are solid at 20 °C.

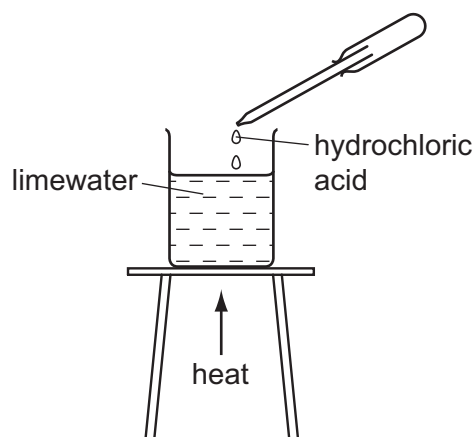
The reactivity of the elements increases down the group.

Which statements about this group of elements and their oxides are correct?

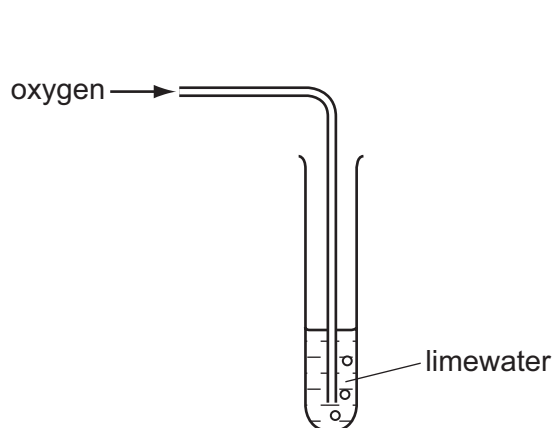
	the elements are in	their oxides are
A	Group I	acidic
B	Group I	basic
C	Group VII	acidic
D	Group VII	basic

26 In which experiment does limewater become milky?

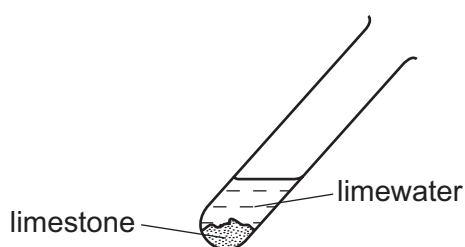
A



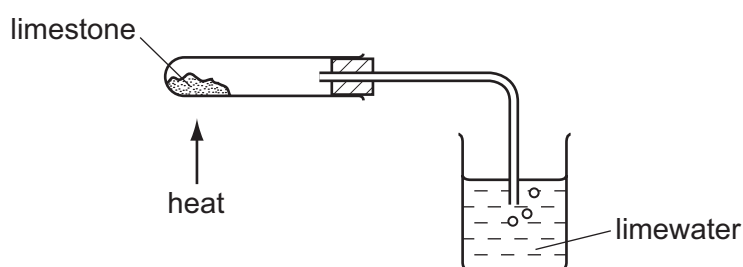
B



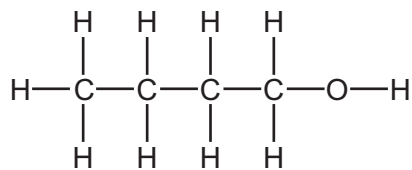
C



D



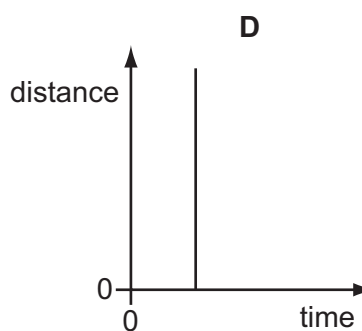
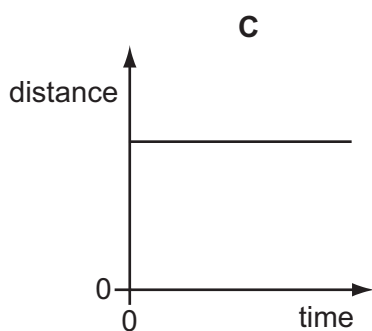
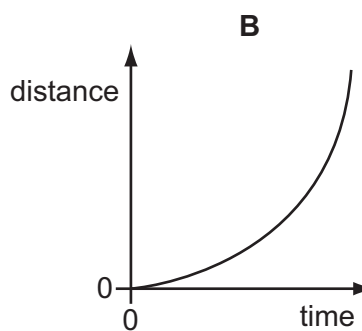
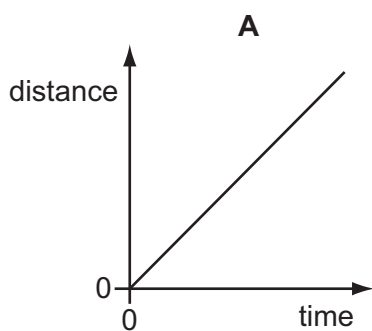
27 The structure of compound P is shown.



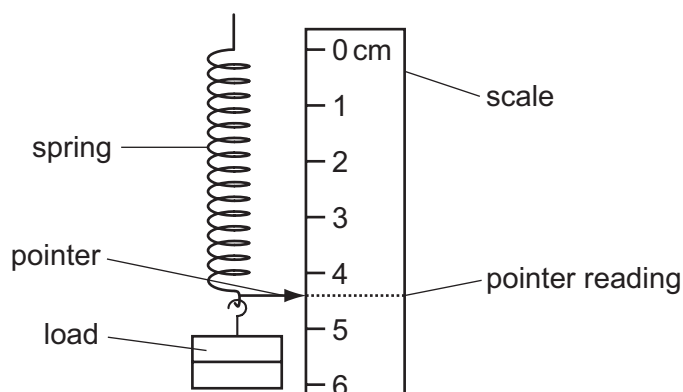
Which type of compound is P?

- A** acid
- B** alcohol
- C** alkane
- D** alkene

28 Which is the distance/time graph for an object moving with constant speed?



29 The diagram shows the arrangement a student uses in an experiment.



She writes down the steps in the order that she follows them, so that she can plot an extension/load graph for the spring.

Which step is **not** correct?

- A Each pointer reading is plotted against the corresponding load.
- B She subtracts the original length of the spring from each pointer reading.
- C The load is added in stages to the lower end of the spring.
- D The reading of the pointer against the scale is recorded for each different load.

30 A student writes an answer.

	Energy is measured in joules.
	Power and work are both
	measured in watts.

Why is this incorrect?

- A Energy is measured in watts.
- B Power is measured in joules.
- C Power is measured in newtons.
- D Work is measured in joules.

- 31 Liquid in a beaker evaporates quickly.

Which row shows what happens to the mass and to the temperature of the liquid in the beaker?

	mass	temperature
A	decreases	decreases
B	decreases	increases
C	increases	decreases
D	increases	increases

- 32 Two identical metal rods are 25 cm long at room temperature (20 °C).

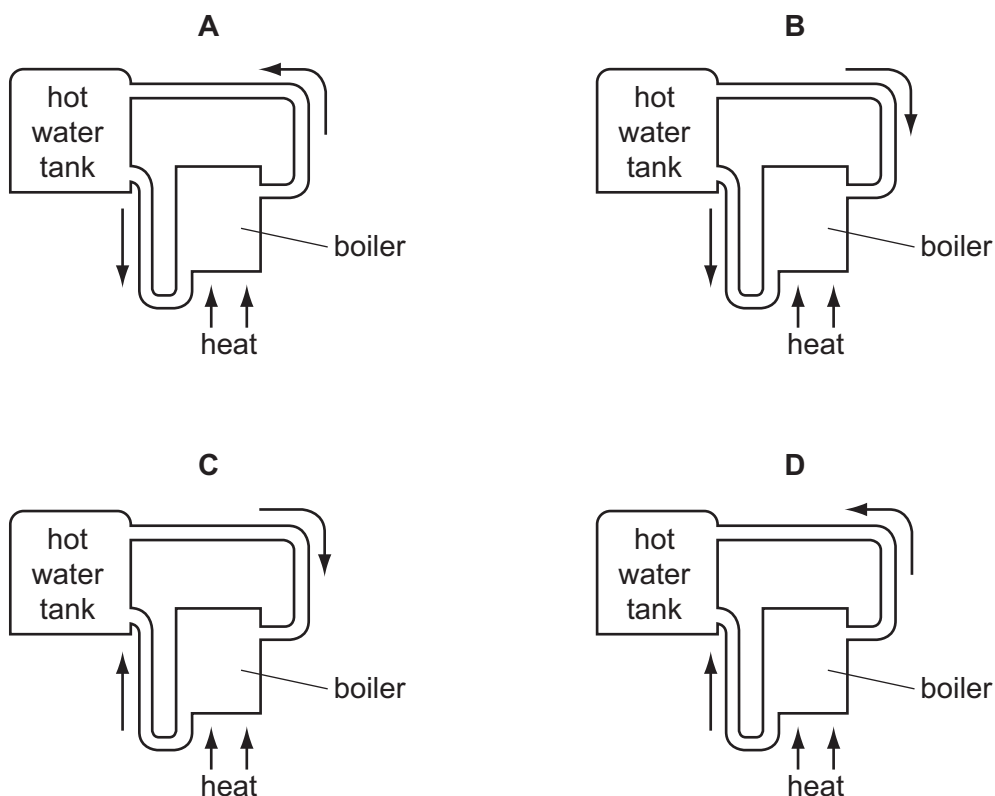
One rod is put into a freezer at a temperature of -18°C . The other rod is put into an oven at a temperature of 200°C . The rods are left for several hours.

Which row shows the new length of each rod?

	length of rod at -18°C	length of rod at 200°C
A	25 cm	25 cm
B	25 cm	more than 25 cm
C	less than 25 cm	25 cm
D	less than 25 cm	more than 25 cm

33 The diagrams show part of a water-heating system which is working by convection.

Which diagram shows the flow of water in the system?



34 Which type of wave is longitudinal?

- A light wave
- B radio wave
- C sound wave
- D water wave

35 A plane mirror forms an image of an object placed in front of it.

Which row describes the image?

	image type	image size
A	real	same size as object
B	real	smaller than object
C	virtual	same size as object
D	virtual	smaller than object

- 36 Red light and violet light have different frequencies and different wavelengths.

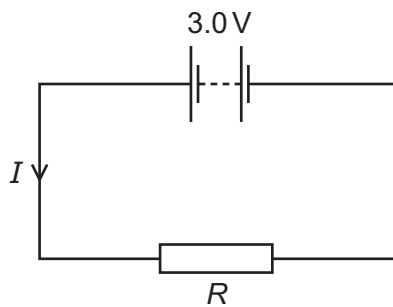
Which color light has the higher frequency and which has the larger wavelength?

	higher frequency	larger wavelength
A	red	red
B	red	violet
C	violet	red
D	violet	violet

- 37 What is the approximate value of the frequency of the highest-pitched sound that can be heard by a young person?

A 20 Hz **B** 200 Hz **C** 2000 Hz **D** 20 000 Hz

- 38 The circuit shows a current I in a resistor of resistance R .



Which row gives possible values of I and of R ?

	I/A	R/Ω
A	1.5	1.5
B	1.5	2.0
C	6.0	2.0
D	4.0	12.0

- 39 Which row shows how lamps are connected in a domestic lighting circuit, and gives an advantage of connecting them in this way?

	how lamps are connected	advantage of connecting them in this way
A	in parallel	they can be switched separately
B	in parallel	they share the voltage
C	in series	they can be switched separately
D	in series	they share the voltage

- 40 An atom of beryllium is represented by ${}^9_4\text{Be}$.

How many neutrons are in the nucleus of this type of beryllium atom?

- A** 4 **B** 5 **C** 9 **D** 13

DATA SHEET
The Periodic Table of the Elements

Group																							
I	II											III	IV	V	VI	VII	0						
												1 H Hydrogen 1											
7 Li Lithium 3	9 Be Beryllium 4																						
23 Na Sodium 11	24 Mg Magnesium 12																						
39 K Potassium 19	40 Ca Calcium 20	45 Sc Scandium 21	48 Ti Titanium 22	51 V Vanadium 23	52 Cr Chromium 24	55 Mn Manganese 25	56 Fe Iron 26	59 Co Cobalt 27	59 Ni Nickel 28	64 Cu Copper 29	65 Zn Zinc 30	70 Ga Gallium 31	73 Ge Germanium 32	75 As Arsenic 33	79 Se Selenium 34	80 Br Bromine 35	84 Kr Krypton 36						
85 Rb Rubidium 37	88 Sr Strontium 38	89 Y Yttrium 39	91 Zr Zirconium 40	93 Nb Niobium 41	96 Mo Molybdenum 42	101 Ru Ruthenium 44	101 Rh Rhodium 45	106 Pd Palladium 46	108 Ag Silver 47	112 Cd Cadmium 48	115 In Indium 49	119 Sn Tin 50	122 Sb Antimony 51	127 Te Tellurium 52	131 I Iodine 53	136 Xe Xenon 54	136 Kr Krypton 36						
133 Cs Caesium 55	137 Ba Barium 56	139 La Lanthanum 57	178 Hf Hafnium 72	181 Ta Tantalum 73	184 W Tungsten 74	186 Re Rhenium 75	190 Os Osmium 76	192 Ir Iridium 77	195 Pt Platinum 78	197 Au Gold 79	201 Hg Mercury 80	204 Tl Thallium 81	207 Pb Lead 82	209 Bi Bismuth 83	210 Po Polonium 84	210 At Astatine 85	222 Rn Radon 86						
87 Fr Francium	88 Ra Radium	89 Ac Actinium																					
58-71 Lanthanoid series																							
90-103 Actinoid series																							
175 Lu Lutetium 71																							
173 Yb Ytterbium 70																							
169 Tm Thulium 69																							
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152 Eu Europium 63																							
150 Sm Samarium 62																							
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141 Pr Praseodymium 59																							
140 Ce Cerium 58																							
238 U Uranium 92																							
235 Pa Protactinium 91																							
232 Th Thorium 90																							
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34 Pt Platinum 78																							
33 Au Gold 79																							
32 Hg Mercury 80																							
31 Tl Thallium 81																							
30 Pb Lead 82																							
29 Bi Bismuth 83																							
28 Po Polonium 84																							
27 At Astatine 85																							
26 Rn Radon 86																							
25 Fr Francium 87																							
24 Ra Radium 88																							
23 Ac Actinium 89																							
22 Th Thorium 90																							
21 Pa Protactinium 91																							
20 U Uranium 92																							
19 Np Neptunium 93																							
18 Pu Plutonium 94																							
17 Am Americium 95																							
16 Cm Curium 96																							
15 Bk Berkelium 97																							
14 Cf Californium 98																							
13 Es Einsteinium 99																							
12 Fm Fermium 100																							
11 Md Mendelevium 101																							
10 No Nobelium 102																							
9 Lr Lawrencium 103																							
8 Lu Lutetium 71																							
7 Yb Ytterbium 70																							
6 Tm Thulium 69																							
5 Er Erbium 68																							
4 Ho Holmium 67																							
3 Dy Dysprosium 66																							
2 Tb Terbium 65																							
1 Gd Gadolinium 64																							
0 Eu Europium 63																							
-1 Sm Samarium 62																							
-2 Nd Neodymium 60																							
-3 Pr Praseodymium 59																							
-4 Ce Cerium 58																							
-5 La Lanthanum 57																							
-6 Hf Hafnium 72																							
-7 Ta Tantalum 73																							
-8 W Tungsten 74																							
-9 Re Rhenium 75																							
-10 Os Osmium 76																							
-11 Ir Iridium 77																							
-12 Pt Platinum 78																							
-13 Au Gold 79																							
-14 Hg Mercury 80																							
-15 Tl Thallium 81																							
-16 Pb Lead 82																							
-17 Bi Bismuth 83																							
-18 Po Polonium 84																							
-19 At Astatine 85																							
-20 Rn Radon 86																							
-21 Fr Francium 87																							
-22 Ra Radium 88																							
-23 Ac Actinium 89																							
-24 Th Thorium 90																							
-25 Pa Protactinium 91																							
-26 U Uranium 92																							
-27 Np Neptunium 93																							
-28 Pu Plutonium 94																							
-29 Am Americium 95																							
-30 Cm Curium 96																							
-31 Bk Berkelium 97																							
-32 Cf Californium 98																							
-33 Es Einsteinium 99																							
-34 Fm Fermium 100																							
-35 Md Mendelevium 101																							
-36 No Nobelium 102																							
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-39 Yb Ytterbium 70																							
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-44 Tb Terbium 65																							
-45 Gd Gadolinium 64																							
-46 Eu Europium 63																							
-47 Sm Samarium 62																							
-48 Nd Neodymium 60																							
-49 Pr Praseodymium 59																							
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-163 Pa Protactinium 91																							
-164 U Uranium 92																							
-165 Np Neptunium 93</																							

*58-71 Lanthanoid series
†90-103 Actinoid series

a	X	b
Key		
a = relative atomic mass	X = atomic symbol	b = proton (atomic) number

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).