

MARK SCHEME for the October/November 2006 question paper

9701 CHEMISTRY

9701/04

Paper 4 (Theory 2), maximum raw mark 60

This mark scheme is published as an aid to teachers and students, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began.

All Examiners are instructed that alternative correct answers and unexpected approaches in candidates' scripts must be given marks that fairly reflect the relevant knowledge and skills demonstrated.

Mark schemes must be read in conjunction with the question papers and the report on the examination.

The grade thresholds for various grades are published in the report on the examination for most IGCSE, GCE Advanced Level and Advanced Subsidiary Level syllabuses.

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	GCE A/AS LEVEL - OCT/NOV 2006	9701	4

- 1 (a) boiling points increase down the group (because of...) (1)
 ...larger van der Waals/intermolecular attractions or bigger induced dipoles (1)
 due to more electrons per molecule (1) [3]
- (b) tetrahedral - clear from diagram (1)
 angles = 109°-110° (1) [2]
- (c) (i) four bonded pairs + 2 lone pairs around Xe (1)
 three lone pairs on at least one F atom (1)
- (ii) square planar (can be read into **very clear** diagram in (i)) (1)
 angles = 90° (1) [4]
- (d) CCl₄ does not react or SiCl₄ does (or read into an equation) (1)
 due to presence of available/low-lying/d-orbitals on Si (1)
- SiCl₄ + 2H₂O → SiO₂ + 4HCl
 (or SiCl₄ + 4H₂O → Si(OH)₄ + 4HCl etc: also allow partial hydrolysis) (1) [3]
- (e) PbCl₄ + 8 Na + 4 C₂H₅Cl → Pb(C₂H₅)₄ + 8 NaCl (1)
 Pb(C₂H₅)₄ = 207 + 4x29 = 323 (1)
 323g needs 8 x 23 = 184g Na
 ∴ 1000g needs 1000 x 184/323 = **569 or 570g** ecf from equ (1)
 (correct ans = (2) marks)
- (alternative method:
 1.0kg of Pb(C₂H₅)₄ is 3.096 moles (1)
 ∴ we need 8 x 3.096 = 24.77 moles of Na, which is **569 or 570g**) (1) [3]

[Total: 15]

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- 2 (a) (i) [one chiral centre only] (1)
- (ii) $C_{13}H_{18}O_2$ (1)
- (iii) $M_r = 206$ ecf (1)
- mass = $0.15 \times (100/1000) \times 206 = 3.1$ g ecf (1)
(correct ans = (2) marks)
- (iv) $n(\text{NaOH}) = 0.1 \times 12/100 = 1.2 \times 10^{-3}$ moles (1)
- $n(\text{A}) = 0.6 \times 10^{-3}$, so $M_r = 0.1/(0.6 \times 10^{-3}) = 167$ (allow 166-170) (1)
(correct ans = (2) marks)
- This fits with **HO₂C-C₆H₄-CO₂H** (which has $M_r = 166$) (1) [7]
- (b) (i) ($K_a =$) $[\text{H}^+][\text{A}^-]/[\text{HA}]$ (1)
- (ii) $[\text{H}^+] = \sqrt{K_a \cdot c} = \sqrt{6.3 \times 10^{-6} \times 0.15} = 9.72 \times 10^{-4}$ (1)
- pH = **3.0** (1)
(correct ans = (2) marks) [3]
- (c) (i) one that **resists/control/maintains** changes in pH (**NOT no** change in pH) (1)
- when **small amounts** of acid/ H^+ (or base/ OH^-) are added. (1)
- (ii) $\text{HPO}_4^{2-} + \text{H}^+ \longrightarrow \text{H}_2\text{PO}_4^-$ (1)
 $\text{H}_2\text{PO}_4^- + \text{OH}^- \longrightarrow \text{HPO}_4^{2-} + \text{H}_2\text{O}$ (1)
- (iii) pH = $\text{p}K_a + \log([\text{base}]/[\text{acid}])$
= $7.2 + \log(.002/.005) = 6.8$ (2)
(correct ans = (2) marks: deduct (1) for each error,
e.g. if ratio is upside down, hence pH = 7.6, answer is worth (1)) [6]

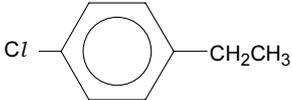
[Total: 16 max 15]

Page 4	Mark Scheme	Syllabus	Paper
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- 3 (a) (i) $2\text{Ca}(\text{NO}_3)_2 \longrightarrow 2\text{CaO} + 4\text{NO}_2 + \text{O}_2$ (or $\times \frac{1}{2}$) (1)
- (ii) (Down the group the nitrates)
 become more stable *or* are more difficult to decompose
or need a higher temperature (to decompose) (1)
 because the radius of **cation/Group II ion/ M^{2+}** increases
or charge density **of the cation** decreases (1)
 thus causing less polarisation/distortion **of the anion/ NO_3^- /nitrate** (1) [4]
- (b) "molar mass" of mixture = $211.6 + 3 \times 12 = 247.6$ (1)
 10 g is thus $10/247.6 = 0.040(4)$ moles (allow ecf for 0.047(3), from $M_r = 211.6$) (1)
 no of moles of gas produced = $0.0404 \times 4 = 0.162$ moles (ecf: 0.189 mol)
 \therefore volume = $0.1616 \times 24 = \mathbf{3.88}$ or $\mathbf{3.9}$ dm^3 (allow ecf for 4.54 dm^3) (1)
 (correct ans = (3) marks)
- (*alternative method:*
 1 mole/247.6g of mixture will produce $4 \times 24 = 96 \text{ dm}^3$ of gas (1)
 \therefore 10g of mixture will produce $96 \times 10/247.6 = \mathbf{3.88}$ or $\mathbf{3.9} \text{ dm}^3$) (1) [3]
- (c) (CO is poisonous...)
 due to complexing/ligand exchange with (Fe of) haemoglobin (1)
 (**NOT** redox involving $\text{Fe}^{2+}/\text{Fe}^{3+}$)
 stopping O_2 being transported around body/in blood/to tissues/from lungs (1) [2]

[Total: 9 max 8]

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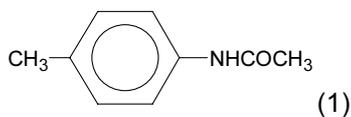
- 4 (a) (i) light or heat [aq or $AlCl_3$ negates] (1)
- (ii) NaOH/KOH/alkali/ OH^- (1)
in alcohol/ethanol + heat [aq negates] (1)
- (iii) $[-CH_2CH(C_6H_5)-]$ [C-C not needed, but C=C is wrong] (1)
- (iv) $CH_2=CHCN$ [C=C is needed here] (1) [5]
- (b) (i) $/OH^-(aq)/NaOH(aq)/aqueous\ alkali/ + heat$ [aq or solution or dil etc. needed] (1)
- (ii) (pale) yellow ppt/crystals (**NOT** orange or orange-yellow) (1)
- (iii) **C/D** is $C_6H_5CO_2Na$ ✓ **D/C** is CHI_3 ✓ (1) + (1) [4]
- (c) (i)
- 
- (1)
- (ii) needs $AlCl_3$ or similar [light or aq negates] (1)
- (iii) (hot) $KMnO_4(aq) + OH^-$ or H^+ [NOT $Cr_2O_7^{2-}$] (1) [3]

[Total: 12]

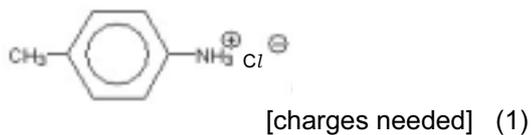
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5 (a) (i) $\text{Br}_2(\text{aq})$ (or solution or in an inert solvent) [light or AlCl_3 etc negates] (1)

(ii) G is



H is



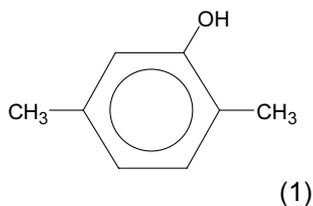
(iii) amide [NOT peptide] (1)

[4]

(b) IV: $\text{H}^+/\text{HCl} + \text{NaNO}_2$ or HNO_2 /nitrous acid (1)

$0^\circ\text{C} \leq T \leq 10^\circ\text{C}$ ["REFLUX" negates] (1)

V:



in $\text{NaOH}(\text{aq})$ (1)

[4]

(c) To increase its solubility in water or to increase binding to food components (1)

due to ionic solvation or more oxygen atoms to H-bond to H_2O /glucose etc (1)

[2]

[Total: 10]